

## Transition Metal Ions and Polyatomic Ions

Write the formulas of the following compounds

1. Copper(I)carbonate

Arsenide

2. Platinum(IV)nitrite

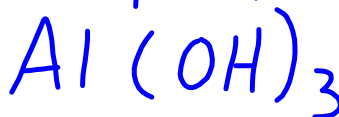
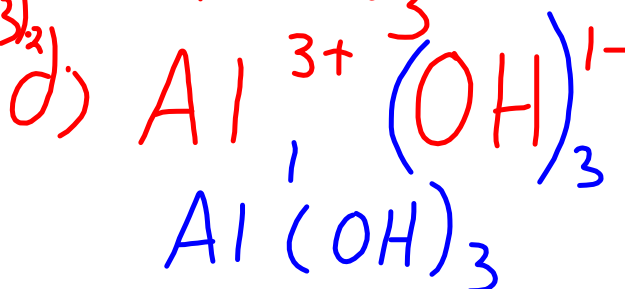
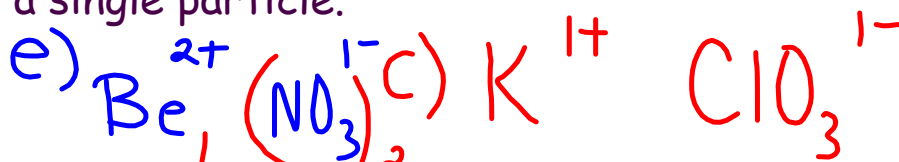
3. Tungsten(V)hydroxide

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#1

Polyatomic ions are a charged group of atoms that chemically act as a single particle.

#3



#4

a) potassium carbonate

b) sodium sulfate

c) aluminum bicarbonate/aluminum hydrogen carbonate

d) silver nitrate

1. Write the formulas for the following compounds.

- |                               |                                |
|-------------------------------|--------------------------------|
| (a) magnesium sulfate _____   | (k) copper(I) chlorate _____   |
| (b) sodium chlorate _____     | (l) calcium sulfate _____      |
| (c) aluminum nitrate _____    | (m) nitric acid _____          |
| (d) potassium hydroxide _____ | (n) carbonic acid _____        |
| (e) lithium phosphate _____   | (o) sulfuric acid _____        |
| (f) calcium carbonate _____   | (p) lead(II) nitrate _____     |
| (g) beryllium sulfate _____   | (q) phosphoric acid _____      |
| (h) sodium bicarbonate _____  | (r) copper(II) hydroxide _____ |
| (i) magnesium hydroxide _____ | (s) iron(II) phosphate _____   |
| (j) aluminum phosphate _____  | (t) calcium chlorate _____     |

2. Write the names for the following compounds.

- |  |  |
|--|--|
| (a) $\text{Li}_2\text{CO}_3$ _____     | (k) $\text{Pb}_3(\text{PO}_4)_2$ _____         |
| (b) $\text{AlHCO}_3$ _____             | (l) $\text{Sn}(\text{ClO}_3)_2$ _____          |
| (c) $\text{Mg}_3(\text{PO}_4)_2$ _____ | (m) $\text{NaOH}$ _____                        |
| (d) $\text{Ca}(\text{NO}_3)_2$ _____   | (n) $\text{H}_3\text{PO}_{4(\text{aq})}$ _____ |
| (e) $\text{K}_2\text{SO}_4$ _____      | (o) $\text{H}_2\text{CO}_{3(\text{aq})}$ _____ |
| (f) $\text{HNO}_{3(\text{aq})}$ _____  | (p) $\text{CuNO}_3$ _____                      |
| (g) $\text{NaNO}_3$ _____              | (q) $\text{H}_2\text{SO}_{4(\text{aq})}$ _____ |
| (h) $\text{Al}(\text{OH})_3$ _____     | (r) $\text{FeSO}_4$ _____                      |
| (i) $\text{CuSO}_4$ _____              | (s) $\text{Ca}(\text{HCO}_3)_2$ _____          |
| (j) $\text{Fe}(\text{ClO})_3$ _____    | (t) $\text{K}_3\text{PO}_4$ _____              |